
Flame Test Lab Metal Ions Answer Key

flame test lab activity key - university of south florida - 10. hold the splint in the flame and record the color of the flame that is produced. 11. using your data, identify the metal ion in your unknown solution. flame test lab activity key note: if chloride compounds are not available, metal nitrate compounds may be substituted. use dilute or approximately 0.1 m solutions unknowns: number the beakers ... **lab- flame tests - rocklin.k12** - distilled water to the test tube if needed using the squirt bottles in your bin. 2. dip a wooden splint into the slurry then insert it into the flame. record your results in the data section. be sure to include such details as whether or not a color is slow in appearing. repeat the ... lab- flame tests **exp 3 flame tests configuration - hcc learning web** - nucleus), energy is released. in the flame test, if this energy has the form of visible light, the flame will produce a color characteristic of the element. different elements have a unique color in its flame which can be used to identify an element. **flame tests abstract - crater high school** - lab partner: avery martin 12/17/12 flame tests abstract: in this flame test experiment, seven different known solutions and three different unknown solutions were tested in a burner. after several tests for each solution it was determined that the **flame test lab handout - annville-cleona school district** - microsoft word - flame test lab handoutcx author: good, brian created date: 12/11/2013 8:22:06 pm ... **lab: flame test - mrs.gillums.science** - flame test lab assessment complete this section after your lab is finished. 1. which of the salt compounds tested was your favorite and why? 2. how many of your color predictions were accurate, not accurate? any major surprises? 3. why do halogens, like chlorine, tend to form compounds with the alkali metals? 4. **cf#5607 flame test kit slk - science notebook** - flame test kit student laboratory kit introduction just as a fingerprint is unique to each person, the color of light emitted by metals heated in a flame is unique to each metal. in this laboratory activity, the characteristic color of light emitted for calcium, copper, lithium, potassium, sodium, and strontium will be observed. chemical concepts

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